

Chemistry 100 (Introduction to Chemistry) Syllabus and Schedule

Basic Course Information

Semester:	Spring 2022	Instructor Name:	Dr. Alto Benedicto
Course Title & #:	Chemistry 100	Email:	alto.benedicto@imperial.edu
CRN #:	20035	Units:	4
Classroom:	Zoom online	Office #:	2779 online
Class Dates:	Feb 14 to Jun 10, 2022	Office Hours:	MTWTh 6:30 am – 7:30 am Zoom
Class Days:	Tutor available MTWThF (see page 5)	Office Phone #:	(760) 355-5751
Class Times:	online	Emergency Contact:	Dept. Secretary (760) 355-6155

Course Description

Elementary principles of general inorganic chemistry with an introduction to organic and biochemistry. Previous science background is recommended but not required. This course is designed for non-science majors and students who need only a one-semester general chemistry course, and also for students entering a paramedical and allied health fields, and industrial applications such as power plants. This course will satisfy the prerequisite for CHEM 200. (CSU)(UC credit limited. See a counselor.) Prerequisite: MATH 091 or MATH 090 with a grade of "C" or better

Student Learning Outcomes

Upon course completion, the successful student will have acquired new skills, knowledge, and or attitudes as demonstrated by being able to:

1. Solve chemical problems using modern atomic theory (ISLO 2, ISLO 4)
2. Perform chemical experiments in a scientific manner, using proper techniques, analysis, and safety equipment. (ISLO 2, ISLO3, ISLO4)

Course Objectives

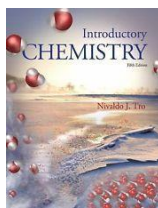
Upon satisfactory completion of the course, students will be able to:

1. calculate English and metric unit conversions and measurements using dimensional analysis.
2. write symbols for elements and know common ionic charges.
3. derive and write formulas and names for chemical compounds.
4. write and balance common chemical equations and identify reaction types.
5. solve stoichiometric problems, including their solutions using dimensional analysis.
6. describe atomic structure including isotopes, periodicity and molecular structure in terms of subatomic particles.
7. identify types of energy and calculate specific heat; identify energy involved in change of state including heat of vaporization and predict behaviors in cooling curves; calculate caloric and nutritional values of various foods.
8. describe gas behavior and solve problems involving the various gas laws.
9. identify the type of intermolecular forces existing between molecules, and its effect on macroscopic property of the substance.
10. calculate solution concentration of various types including dilutions.

11. define the three basic concepts (Arrhenius, Bronsted-Lowry and Lewis) of acids and bases and perform titration experiments and calculate pH.
12. use Le Chatelier's Principle to predict the shift in the direction of the reactants/products
13. determine the oxidant/reductant and balance redox equations.
14. describe nuclear processes and write nuclear equations using the subatomic particles involved and identify health factors and risks involved.

Textbooks & Other Resources or Links

1. *Introductory Chemistry*, by Nivaldo J. Tro (Custom Edition for IVC. Prentice-Hall Publishing, 5th Ed, ISBN: 1269713876)



NOTE: There is a SHORTENED ONLINE version at

[https://chem.libretexts.org/Bookshelves/Introductory_Chemistry/Map%3A_Introductory_Chemistry_\(Tro\)](https://chem.libretexts.org/Bookshelves/Introductory_Chemistry/Map%3A_Introductory_Chemistry_(Tro))

- ~~2. Chemistry 100 Laboratory Manual available at **IVC Chemistry/STEM Club** (\$15) **Free access to LABSTER simulations.**~~
- ~~3. Eight (8) Scantron Sheets Form No. 882-E (submitted on the second day of class) and pencil~~
- ~~4. safety goggles (\$5 – \$10; needed on second class day), **non-programmable scientific calculator** (\$15 - \$25), close-toed shoes~~
5. registration with <https://achieve.macmillanlearning.com/courses> for online HW (\$42) – requires credit card. Course ID: **k4pxte (for CRN 20035)** . For help: <https://macmillan.force.com/macmillanlearning/s/article/Achieve-Access-Code-FAQ> .
6. free access to “Online Tutoring” seven days a week (online tutoring with a live person in California) via Canvas

Course Requirements and Instructional Methods

- ~~1. Attendance and remaining during the entire class period is mandatory for Chem 100 Lab Classes. A Lab roll call will be initiated by the instructor within the first 5 minutes of Lab class. If you are sent out during class (e.g., failure to obey safety rules such as wearing Safety Goggles, etc.), you will be marked absent.~~
2. There are **no make-up Exams or Lab Classes**. A score of **zero (0)** will be recorded unless the absence is attributed to representation of official college functions. It is the student’s responsibility to show proof of such function **prior** to the date of the absence.
3. During Exam, the only things allowed are: **pencil, nonprogrammable calculator, and I.D.** You will be supplied with a Periodic Table and a Scantron. ~~You may use the Exam Questionnaire as scratch paper. The Exam Questionnaire, Periodic Table, and Scantron are to be submitted at the end of the Exam.~~ **Possession of electronic devices (phones, ipad, programmable calculator, etc.) during Exam is considered cheating and will be dealt with according to IVC policy.**
4. Each student is REQUIRED to ~~buy the Chem 100 Lab Manual~~ and to **sign up for online homework (HW) no later than the second day of class**. Personal laptop is highly encouraged for online HW during Lab Class.
5. **Due dates for Online HWs are found in the Class Schedule of Topics (see last page)**. For technical assistance beyond the instructor, call Macmillan Technical Support at 1-800-936-6899. Also, there’s online tutoring with a live person in **Online Tutoring** (embedded inside Canvas).

6. ~~Prior to start of Lab Class, read the relevant experiment and answer any Pre-Lab Questions. **Pre-Lab Questions sheet should be torn from the Lab Manual and submitted to the Instructor within two (2) minutes from start of Lab Class to gain full points.** So tear out the relevant Pre-Lab sheets before coming to class, and don't be late!!!~~
7. ~~Before leaving the Lab Class, make sure the **instructor has signed** your Lab Data Sheet. Data should be recorded in **ink**. Cross-out mistakes with a single strike-through line. **Data Sheets and Post-Lab Questions are to be submitted within the first two minutes of the next time Lab meeting.**~~
8. ~~Lab clean-ups are done 15 minutes before the end of lab. A **wet towel** should be used to wipe the lab bench in order to gain full points. Make sure sink and work area is clean. Points will be deducted to the entire class if the common work areas (fume hood, analytical balances) are dirty.~~
9. There is no bonus work available. Kindly seek assistance immediately to clarify any questions.
10. If this is a Hybrid section, with the lecture discussion being done online, you must have access to a computer and an Internet connection. No other special technical skills are needed other than knowledge on how to use Canvas.

Out of Class Assignments: The Department of Education policy states that one (1) credit hour is the amount of student work that reasonably approximates not less than one hour of class time and two (2) hours of out-of-class time per week over the span of a semester. WASC has adopted a similar requirement.

Course Grading Based on Course Objectives

Assessment Type	How many	Total Points
Lecture Exams	6 @ 50	300 pts
Lecture Final Exam	1 @ 150	150 pts
Online Homework	11 @ 20 6 @ 25	370 pts
Labster Simulations and PhET Simulations	11 @ 15; 1 @ 5 1 @ 20	190 pts
Lab Exam	1 @ 50	50 pts
Canvas Discussion	5 @ 5	25 pts

OVERALL POINTS = 1,085 pts

Grading Scale Percentage	Letter Grade
85.00% to 100 %	A
75.00% to 84.99%	B
60.00% to 74.99%	C
50.00% to 59.99%	D

Attendance

- A student who fails to attend the first meeting of a class or does not complete the first mandatory activity of an online class will be dropped by the instructor as of the first official meeting of that class. Should readmission be desired, the student's status will be the same as that of any other student who desires to add a class. It is the student's responsibility to drop or officially withdraw from the class. See General Catalog for details.

- Regular attendance in all classes is expected of all students. **A student whose continuous, unexcused absences exceed the number of hours the class is scheduled to meet per week may be dropped.** For online courses, students who fail to complete required activities for two consecutive weeks may be considered to have excessive absences and may be dropped.
- Absences attributed to the representation of the college at officially approved events (conferences, contests, and field trips) will be counted as ‘excused’ absences.
- **Absences during Lab Classes, or leaving during Lab Classes** automatically result in a **grade of zero (0) for the Lab Experiment.**

Classroom Etiquette

- Electronic Devices: Cell phones and electronic devices must be turned off and put away during class, unless otherwise directed by the instructor.
- Food and Drink are prohibited in all classrooms. Water bottles with lids/caps are the only exception. Additional restrictions will apply in labs. Please comply as directed by the instructor.
- Disruptive Students: Students who disrupt or interfere with a class may be sent out of the room and told to meet with the Campus Disciplinary Officer before returning to continue with coursework. Disciplinary procedures will be followed as outlined in the [General Catalog](#).
- Children in the classroom: Due to college rules and state laws, no one who is not enrolled in the class may attend, including children.

Online Netiquette

- What is netiquette? Netiquette is internet manners, online etiquette, and digital etiquette all rolled into one word. Basically, netiquette is a set of rules for behaving properly online.
- Students are to comply with the following rules of netiquette: (1) identify yourself, (2) include a subject line, (3) avoid sarcasm, (4) respect others’ opinions and privacy, (5) acknowledge and return messages promptly, (6) copy with caution, (7) do not spam or junk mail, (8) be concise, (9) use appropriate language, (10) use appropriate emoticons (emotional icons) to help convey meaning, and (11) use appropriate intensifiers to help convey meaning [do not use ALL CAPS or multiple exclamation marks (!!!!)].

Academic Honesty

Academic honesty in the advancement of knowledge requires that all students and instructors respect the integrity of one another’s work and recognize the important of acknowledging and safeguarding intellectual property.

There are many different forms of academic dishonesty. The following kinds of honesty violations and their definitions are not meant to be exhaustive. Rather, they are intended to serve as examples of unacceptable academic conduct.

- Plagiarism is taking and presenting as one’s own the writings or ideas of others, without citing the source. You should understand the concept of plagiarism and keep it in mind when taking exams and preparing written materials. If you do not understand how to “cite a source” correctly, you must ask for help.
- Cheating is defined as fraud, deceit, or dishonesty in an academic assignment, or using or attempting to use materials, or assisting others in using materials that are prohibited or inappropriate in the context of the academic assignment in question.

Anyone caught cheating or plagiarizing will receive a zero (0) on the exam or assignment, and the instructor may report the incident to the Campus Disciplinary Officer, who may place related documentation in a file. Repeated acts of cheating may result in an F in the course and/or disciplinary action. Please refer to the [General Catalog](#) for more

information on academic dishonesty or other misconduct. Acts of cheating include, but are not limited to, the following: (a) plagiarism; (b) copying or attempting to copy from others during an examination or on an assignment; (c) communicating test information with another person during an examination; (d) allowing others to do an assignment or portion of an assignment; (e) using a commercial term paper service.

Additional Student Services

Imperial Valley College offers various services in support of student success. The following are some of the services available for students. Please speak to your instructor about additional services which may be available.

- **Learning Services.** There are several learning labs on campus to assist students through the use of computers and tutors. Please consult your [Campus Map](#) for the [Math Lab](#); [Reading, Writing & Language Labs](#); and the [Study Skills Center](#).
- **Library Services.** There is more to our library than just books. You have access to tutors in the [Study Skills Center](#), study rooms for small groups, and online access to a wealth of resources.

Disabled Student Programs and Services (DSPS)

Any student with a documented disability who may need educational accommodations should notify the instructor or the [Disabled Student Programs and Services](#) (DSP&S) office as soon as possible. The DSP&S office is located in Building 2100, telephone 760-355-6313. Please contact them if you feel you need to be evaluated for educational accommodations.

Student Counseling and Health Services

Students have counseling and health services available, provided by the pre-paid StudentHealth Fee.

- **Student Health Center.** A Student Health Nurse is available on campus. In addition, Pioneers Memorial Healthcare District provide basic health services for students, such as first aid and care for minor illnesses. Contact the IVC [Student Health Center](#) at 760-355-6128 in Room 1536 for more information.
- **Mental Health Counseling Services.** Short-term individual, couples, family, and group therapy are provided to currently enrolled students. Contact the IVC [Mental Health Counseling Services](#) at 760-355-6196 in Room 2109 for more information.

Student Rights and Responsibilities

Students have the right to experience a positive learning environment and to due process of law. For more information regarding student rights and responsibilities, please refer to the IVC [General Catalog](#).

Information Literacy

Imperial Valley College is dedicated to helping students skillfully discover, evaluate, and use information from all sources. The IVC [Library Department](#) provides numerous [Information Literacy Tutorials](#) to assist students in this endeavor.

TUTORING RESOURCES:

- 1) Our Class Tutor _____ hours are (MWF: _____)(TThS: _____) (Zoom ID 930 5535 9930—ask for _____)
- 2) My Tutoring/Office Hours: MTWTh 6:30 am – 7:30 am. (Zoom link in Canvas Announcement)
- 3) **Online Tutoring** in left margin of Canvas seven days a week allows you to access live tutoring from State of California

Anticipated Class Schedule/Calendar. *Tentative, subject to change without prior notice*****

Day	DATE	CHAPTER READINGS (<i>Watch all of my Zoom recordings AND YouTube videos</i>)	<i>Achieve Online Homework and Canvas Discussion due at 11:55 pm Sat</i>	LABORATORY via LABSTER Simulation (<i>all due Jun 8</i>), <i>however do them at suggested day below if you can</i>
1	Feb 14 – Feb 19	Orientation 6:30 am Tues Feb 15 (Zoom Meet ID will be emailed to Registered & WaitListed on Feb 15) Ch 1: Chemical World	HW 1 due (in Achieve—by Macmillan Learning); <i>Canvas Discussion #1 due</i>	IVC safety quiz due (in Webstar; see how to access by watching Zoom); Labster 1: Lab Safety ‘due’ (in Canvas)
2	Feb 21 – Feb 26	Ch 2: Measurement	HW 2 due	Labster 2: Matter and its Phase Changes (in Canvas)
3	Feb 28 – Mar 5	Ch 3: Matter and Energy	HW 3 due	Labster 3: Periodic Table of Elements
4	Mar 7 – Mar 12	Ch 4: Atoms and Elements	HW 4 due <i>PhET Sim: Build An Atom due (in Canvas)</i>	Lecture Exam 1 on Thurs 6 pm - 7:40 pm (covers Ch 1, 2, 3)
5	Mar 14 – Mar 19	Ch 5: Molecules and Compounds	HW 5 due	Labster 4: Atomic Structure
6	Mar 21 – Mar 26	Ch 6: Chemical Composition	HW 6 due <i>Canvas Discussion #2 due</i>	<i>Canvas Discussion #2 due</i>
7	Mar 28 – Apr 2	Ch 7: Chemical Reactions	HW 7 due	Lecture Exam 2 on Thurs 6 pm - 7:40 pm (covers Ch 4, 5, 6)
8	Apr 4 – Apr 9	Ch 8: Quantities in Chemical Reactions	HW 8 due	Labster 5: Stoichiometric Calculations
9	Apr 11 – Apr 16	Ch 9: Electrons in Atoms and the Periodic Table	HW 9 due	Lecture Exam 3 on Thurs 6 pm - 7:40 pm (covers Ch 7,8)
10	Apr 18 – Apr 23	Ch 10: Chemical Bonding	HW 10 due <i>Canvas Discussion #3 due</i>	Labster 6: Bohr & Quantum Models Labster 7: Ionic and Covalent Bonds
11	Apr 25 – Apr 30	Ch 11: Gases	HW 11 due	Lecture Exam 4 on Thurs 6 pm - 7:40 pm (covers Ch 9, 10)
12	May 2 – May 7	Ch 12: Liquids, Solids, and Intermolecular Forces	HW 12 due	Labster 8: Ideal Gas
13	May 9 – May 14	Ch 13: Solutions	HW 13 due <i>Canvas Discussion #4 due</i>	Labster 9: Solution Preparation
14	May 16 – May 21	Ch 14: Acids and Bases	HW 14 due	Lecture Exam 5 on Thurs 6 pm - 7:40 pm (covers Ch 11,12,13) Labster 10: Acids and Bases
15	May 23 – May 28	Ch 15: Chem Equilibrium	HW 15 due	Labster 11: Titration Labster 12: Equilibrium
16	May 30 – Jun 4	Ch 16: Redox Reaction Lecture on Laboratory Techniques for Lab Finals	HW 16 due <i>Canvas Discussion #5 due</i>	Lecture Exam 6 on Thurs 6 pm - 7:40 pm (covers Ch 14, 15, 16)
17	Jun 6 – Jun 9	Ch 17: Radioactivity & Nuclear Chem	HW 17 due on Wed Jun 8 at 11:55 pm	LAB FINAL EXAM at 6-7 pm on Tues LEC FINAL EXAM at 6-8:10 pm on Th

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