

Chemistry 200 (General Inorganic Chemistry I) Syllabus and Schedule

Basic Course Information

Semester:	Summer 2023	Instructor Name:	Dr. Alto Benedicto
Course Title & #:	Chemistry 200	Email:	alto.benedicto@imperial.edu
CRN #:	30174	Units:	5
Classroom:	Zoom online	Office #:	2779 online
Class Dates:	Jun 20 to Jul 27, 2023	Office Hours:	Zoom by appointment
Class Days:	Tutoring available (see page 5)	Office Phone #:	(760) 355-5751
Class Times:	online	Emergency Contact:	Dept. Secretary (760) 355-6155

Course Description

Basic principles and calculations of chemistry with emphasis on stoichiometry and dimension analysis applied to various problem types. Fundamental principles and theory of atomic and molecular structure as related to bonding and molecular geometry. Study of kinetic molecular theory, the first law of thermodynamics, periodic relationships of the elements, physical states of matter, solution chemistry, and oxidation-reduction. The laboratory is closely related to lecture topics and includes methods of classical experimentation as well as certain instrumental analysis. (C-ID CHEM 110) (CSU, UC).

Course Prerequisite(s) and/or Corequisite(s)

Chem 100.

Student Learning Outcomes

Upon course completion, the successful student will have acquired new skills, knowledge, and or attitudes as demonstrated by being able to:

1. Solve chemical problems using modern atomic theory (ISLO 2, ISLO 4)
2. Perform chemical experiments in a scientific manner, using proper techniques, analysis, and safety equipment. (ISLO 2, ISLO3, ISLO4)

Course Objectives

Upon satisfactory completion of the course, students will be able to:

1. Student will demonstrate ability to perform dimensional analysis calculations as they relate to problems involving percent composition and density.
2. Student will write chemical formulas, name inorganic compounds, and demonstrate a knowledge of basic atomic theory
3. Student will relate chemical equations and stoichiometry as they apply to the mole concept, molarity, and acid-base titrations. Student will derive formulas from percent composition.

4. Student will identify the basic types of chemical reactions including precipitation, neutralization, and oxidation-reduction.
5. Student will demonstrate knowledge of atomic structure and quantum mechanics and apply these concepts to the study of periodic properties of the elements.
6. Student will relate the general concepts of atomic structure to a study of ionic bonding.
7. Student will relate the general concepts of covalent bonding and molecular structure.
8. Student will demonstrate the first law of thermodynamics both in theoretical and practical contexts and apply the theory to the solution of Hess' Law.
9. Student will manipulate the various gas laws in both theory and practice to solve mathematical problems relating to the behavior of both ideal and non-ideal gases.
10. Student will describe the general properties of liquids and solids including intermolecular attractions and phase changes.
11. Student will relate the general properties of solutions and employ knowledge of concentration to explain colligative properties. Student will investigate the phenomenon of vapor pressure.
12. Student will demonstrate knowledge of computer-assisted methods of data acquisition, analysis and presentation.

Textbooks & Other Resources or Links

1. **Two sites used by IVC for our textbooks:**

- a. <https://chem.libretexts.org>
 - i. For the textbook: click [https://chem.libretexts.org/Bookshelves/General_Chemistry/Chemistry_\(OpenSTAX\)](https://chem.libretexts.org/Bookshelves/General_Chemistry/Chemistry_(OpenSTAX))
 - ii. For exercises: click https://chem.libretexts.org/Bookshelves/General_Chemistry/Exercises%3A_General_Chemistry/Exercises%3A_OpenStax
- b. Chemistry, by Paul Flowers, Klaus Theopold, Richard Langley, Stephen F. Austin, and William R. Robinson (OpenStax, 2015, ISBN: 1-938168-39-9)
 - i. For this textbook: click <https://openstax.org/details/books/chemistry>

- ~~2. *Chemistry in the Laboratory*, by James M. Postma, Julian Roberts, and J. Leland Hollenberg (W.H. Freeman and Company, 7th Ed, ISBN: 1-4292-1954-8)~~
- ~~3. Chemistry 200 Supplementary Laboratory Manual available at **IVC Chemistry/STEM Club** (\$15)~~
- ~~4. Eight (8) Scantron Sheets Form No. 889-E (submitted on the second day of class) and pencil~~
5. safety goggles (\$5 – \$10; needed on second class day), **non-programmable scientific calculator** (\$15 - \$25), close-toed shoes.
6. registration with Macmillan Learning Achieve for online HW (\$50) – requires credit card.
You can register by going to our course in Canvas, and then clicking Macmillan Learning (located on left margin) while INSIDE our Canvas course, and follow instructions.

Course Requirements and Instructional Methods

1. Attendance and remaining during the entire class period is mandatory for Chem 200 Lab Classes. A Lab roll call will be initiated by the instructor within the first 5 minutes of Lab class. If you are sent out during class (e.g., failure to obey safety rules such as wearing Safety Goggles, etc.), you will be marked absent for that Lab, and will garner zero points for the experiment.

- There are **no make-up Exams or Lab Classes**. A score of **zero (0)** will be recorded unless the absence is attributed to representation of official college functions. It is the student's responsibility to show proof of such function **prior** to the date of absence.
- During the Exam, the only things allowed are: **pencil, nonprogrammable calculator, and I.D.** You will be supplied with a Periodic Table and a Scantron. You may use the Exam Questionnaire as scratch paper. The Exam Questionnaire, Periodic Table, and Scantron are to be submitted at the end of the Exam. **Possession of electronic devices (phones, ipod, programmable calculator, etc.) during Exam is considered cheating** and will be dealt with according to IVC policy.
- Each student is **REQUIRED** to ~~buy the Chem 200 Supplement Lab Manual~~ and to **sign up for online HW no later than the second day of class**. Personal laptop is highly encouraged for online HW during Lab Class.
- Due dates for Online HWs are found in the Class Schedule of Topics (see last page)**. For technical assistance beyond the instructor, call Macmillan Technical Support at 1-800-936-6899. Also, there's online tutoring with a live person in **Online Tutoring** (embedded inside Canvas).
- ~~Prior to the start of Lab Class, read the relevant experiment and answer any Pre-Lab Questions. Pre-Lab Questions sheet should be torn from the Lab Manual and submitted to the instructor within two (2) minutes from start of Lab Class to gain full points.~~ So, tear out the relevant Pre-Lab sheets before coming to class, and don't be late!!!
- ~~Before leaving the Lab Class, make sure the instructor has signed your Lab Data Sheet. Data should be recorded in ink. Cross-out mistakes with a single strike-through line. Data Sheets and Post-Lab Questions are to be submitted within the first two minutes of the next time Lab meeting.~~
- Lab clean-ups are done 15 minutes before the end of lab. A **wet towel** should be used to wipe the lab bench in order to gain full points. Make sure the sink and work area is clean. Points will be deducted to the entire class if the common work areas (fume hood, analytical balances) are dirty.
- There is no bonus work available. Kindly seek assistance immediately to clarify any questions.
- If this is an Hybrid section, with the lecture discussion being done online, you must have access to a computer and an Internet connection. ~~No other special technical skills are needed other than knowledge on how to use Canvas.~~

Out of Class Assignments: The Department of Education policy states that one (1) credit hour is the amount of student work that reasonably approximates not less than one hour of class time and two (2) hours of out-of-class time per week over the span of a semester. WASC has adopted a similar requirement.

Course Grading Based on Course Objectives

Assessment Type	How many	Total Points
Lecture Exams	5 @ 60	300 pts
Lecture Final Exam	1 @ 200	200 pts
Online Homework (in Achieve)	11 @ 15 3 @ 20; 1 @ 30	255 pts
Labster Simulations and PhET Simulations	12 @ 12 1 @ 6, 1 @ 20	170 pts
Lab Exam and Discussion	1 @ 50+25	75 pts

OVERALL POINTS = 1,000 pts

Grading Scale Percentage	Letter Grade
85.00% to 100 %	A

Grading Scale Percentage	Letter Grade
75.00% to 84.99%	B
60.00% to 74.99%	C
50.00% to 59.99%	D

Course Policies

- A student who fails to attend the first meeting of a class or does not complete the first mandatory activity of an online class will be dropped by the instructor as of the first official meeting of that class. Should readmission be desired, the student's status will be the same as that of any other student who desires to add a class. It is the student's responsibility to drop or officially withdraw from the class. See General Catalog for details.
- Regular attendance in all classes is expected of all students. **A student whose continuous, unexcused absences exceed the number of hours the class is scheduled to meet per week may be dropped.** For online courses, students who fail to complete required activities for two consecutive weeks may be considered to have excessive absences and may be dropped.
- Absences attributed to the representation of the college at officially approved events (conferences, contests, and field trips) will be counted as 'excused' absences.
- **Absences during Lab Classes or leaving during Lab Classes** automatically result in a **grade of zero (0) for the Lab Experiment.**

Academic Honesty

- IVC values critical thinking and communication skills and considers academic integrity essential to learning. Using AI tools as a replacement for your own thinking, writing, or quantitative reasoning goes against both our mission and academic honesty policy and will be considered academic dishonesty, or plagiarism unless you have been instructed to do so by your instructor. In case of any uncertainty regarding the ethical use of AI tools, students are encouraged to reach out to their instructors for clarification.

Anyone caught cheating or plagiarizing will receive a zero (0) on the exam or assignment, and the instructor may report the incident to the Campus Disciplinary Officer, who may place related documentation in a file. Repeated acts of cheating may result in an F in the course and/or disciplinary action. Please refer to the [General Catalog](#) for more information on academic dishonesty or other misconduct. Acts of cheating include, but are not limited to, the following: (a) plagiarism; (b) copying or attempting to copy from others during an examination or on an assignment; (c) communicating test information with another person during an examination; (d) allowing others to do an assignment or portion of an assignment; (e) using a commercial term paper service.

IVC Student Resources

IVC wants you to be successful in all aspects of your education. For help, resources, services, and an explanation of policies, visit <http://www.imperial.edu/studentresources> or click the heart icon in Canvas.

TUTORING RESOURCES:

- 1) Our Class Tutor _____ hours are (MWF: _____)(TThS: _____) (Zoom ID 930 5535 9930—ask for _____)
- 2) My Tutoring/Office Hours: by appointment ??? (Zoom link in Canvas Announcement)
- 3) **Online Tutoring** in left margin of Canvas seven days a week allows you to access live tutoring from State of California

Anticipated Class Schedule/Calendar

Day	DATE	CHAPTER READINGS	Laboratory LABSTER (all due July 26); Macmillan Achieve HW due at 11:55 pm
1	June 20 (T)	Orientation 7:00 pm WED Jun 21 (Zoom Meet ID will be emailed to Registered & WaitListed on Jun 21) Chap 1: The Chemical World (Classification of Matter/Changes; Measurement—Sig Figs, Dimensional Analysis, Accuracy)	Labster 1: Lab Safety 'due' (in Canvas) <i>Canvas Discussion 1 due</i> HW 1 due next day
2	Jun 21 (W)	Chap 2: Atoms, Molecules, Ions (Atomic Theory; Atomic Number, Mass Number, Isotopes; Nomenclature of Hydrates)	Labster 2: Matter and its Phase Changes (in Canvas) HW 2 due next day <i>PhET Sim: Build An Atom due Thurs Jun 22 (in Canvas)</i>
3	Jun 22 (Th)	Chap 3: Composition of Substances and Solutions (Moles, Molar Mass, % Composition, Empirical & Molec Formula, Molarity)	Labster 3: Periodic Table of Elements HW 3 due next day
4	Jun 23 (F)	Chap 4: Stoichiometry (Balanced Eqn, Classifying Rxns, Yields)	Labster 4: Atomic Structure HW 4 due next day
5	Jun 26 (M)	Lecture Exam 1 (Chap 1,2,3,4) 7 pm - 8:40 pm	<i>Canvas Discussion 2 due</i>
6	Jun 27 (T)	Chap 6: Electronic Structure and Periodic Properties (Bohr Model, QM model, Quantum Numbers, Electron Configurations, Orbital Diagrams, Periodic Trends)	Labster 5: Stoichiometric Calculations HW 6 due next day
7	Jun 28 (W)	Chap 7: Chemical Bonding I (exceptions to Lewis structure, Ionic/Covalent/Metallic Bonding, Formal Charges, Resonance, Molecular Geometry, Polarity)	Labster 6: Bohr and Quantum Model HW 7 due next day
8	Jun 29 (R)	Chap 8: Chemical Bonding II: (Valence Bond vs Molecular Orbital Theory)	Labster 7: Ionic and Covalent Bonds HW 8 due next day
9	July 3 (M)	Lecture Exam 2 (Chap 6,7,8) 7 pm - 8:40 pm Chap 9: Gases (Ideal Gas Law, Real Gases, Effusion vs Diffusion, Stoichiometry)	HW 9 due next day
10	July 5 (W)	Chap 10: Liquids and Solids (Intermolecular Forces, Phases Diagrams, Lattice Structures, Calorimetry)	Labster 8: Ideal Gas Law HW 10 due next day
11	July 6 (Th)	Chap 11: Solutions and Colloids: (Solubility; Colligative Properties; Colloids, molality)	Labster 9: Solution Preparation HW 11 due next day
12	July 10 (M)	Lecture Exam 3 (Chap 9,10,11, stoichiometry) 7 pm - 8:40 pm	<i>Canvas Discussion 3 due</i>
13	July 11 (T)	Chap 13: Chemical Equilibrium (Equilibrium Constant, Le Chatelier's Principle, ICE box)	Labster 12: Equilibrium HW 13 due next day

Day	DATE	CHAPTER READINGS	Laboratory LABSTER (all due July 26); Macmillan Achieve HW due at 11:55 pm
14	July 12 (W)	Chap 14: Acid-Base Equilibria (definitions, pH, weak acids, pKa, polyprotic acid, buffers)	Labster 10: Acids and Bases
15	Jul 13 (Th)	Con't of Chap 14: (Titration; hydrolysis of salt solutions)	Labster 11: Titration HW 14a due next day
16	Jul 17 (M)	Lecture Exam 4 (Chap 13,14) 7 pm - 8:40 pm	<i>Canvas Discussion 4 due</i>
17	July 18 (T)	Chap 15: Solubility Equilibria (Ksp, complex ion formation, common ion effect)	Labster 13: Advanced Acid and Base HW 15 due next day
18	July 19 (W)	Chap 5: Thermochemistry (Calorimetry, Hess Law, Std Enthalpy of Formation)	HW 5 due next day
19	Jul 20 (Th)	Ch 16: Thermodynamics 1 st Law (Bond Dissociation Energy, Born-Haber cycle) NOTE: Omit Second and Third Law for next course	HW (5 and 16) due next day
20	Jul 24 (M)	Lecture Exam 5 (Chap 15,5,16) 7 pm - 8:40 pm	<i>Canvas Discussion 5 due</i>
21	July 25 (T)	Lecture on Lab Techniques	HW 14b due next day
22	July 26 (W)	Review for Finals	Lab Exam (Lab Techniques) 7 pm - 8 pm (1 hour)
23	Jul 27 (Th)	(Finals is in Macmillan Achieve)	LEC FINAL EXAM (7 pm - 9:10 pm) (2 hr 10 min)

*****Tentative, subject to change without prior notice*****